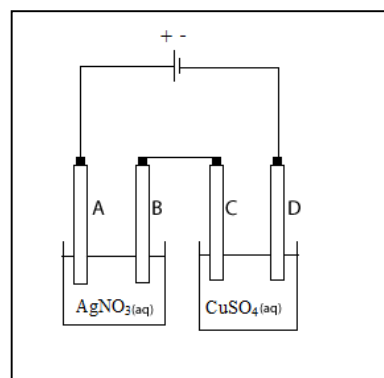


Friday quiz
electrolysis – worksheet 12

In an attempt to produce silver and copper metals two electrolytic cells are setup in series as shown on the right.

- i. Give the products formed (if any) at each electrode.

A. _____
B. _____
C. _____
D. _____



- ii. Assuming 100% efficiency calculate the mol of silver metal and copper metals deposited if an electric charge of 2.50 faradays is passed through the circuit.
A charge of 2.5 faradays is equivalent to 2.5 mol of electrons.

- iii. What will be the change in mass at each electrode? Explain your answer and calculate the mass change.

- iv. Indicate the direction of electron flow.

- v. If a current of 1.25 A runs through the circuit calculate the time, in hours, required to deposit 0.136 Kg of copper.